

Chapter 1 & 3 review

Name _____

Define the following:

1. Control -
2. Hypothesis -
3. Conclusion -
4. Theory –
5. Law -
6. Model –
7. Independent Variable -
8. Dependent Variable -
9. Matter –
10. Element –
11. Compound-
12. Homogeneous Mixture -
13. Heterogeneous Mixture –
14. Distillation -
15. Filtration –
16. Chromatography -
17. Crystallization -
18. Reactants -
19. Products -
20. Chemical Equation -
21. Law of Conservation of Mass -
23. A 3.50-g sample of pure copper reacts with chlorine to form 8.76 g of a copper chloride compound. How many grams of chlorine are used in this reaction? (show work and give units)

22. Identify the following as quantitative or qualitative data...

Color -

Volume -

Pressure -

Shape -

Temperature -

Odor-

24. Identify the following as a chemical or physical property...

Remains unchanged in oxygen -

Photosynthesis -

Liquid at room temperature -

Burning coal –

Blue -

Snapping a branch off a tree –

Conducts electricity -

Water boiling –

Changes colors when heated -

Chrome tarnishing –

Grinding cocoa beans –

Baking a cake -

25. Identify the following as heterogeneous or homogeneous mixtures

Blood -

Gasoline -

Limestone -

Baking soda -

Polluted air -

Top soil -

Glass -

Grape juice -

Lotion -

Milk -

26. Identify the following as an element or a compound

Copper-

Carbon Monoxide -

Vinegar -

Bronze -

Oxygen -

Plutonium -

Chapter 1 & 3 review

KEY

Define the following: **See you book!**

1. Control -

2. Hypothesis -

3. Conclusion -

4. Theory –

5. Law -

6. Model –

7. Independent Variable -

8. Dependent Variable -

9. Matter –

10. Element –

11. Compound-

12. Homogeneous Mixture -

13. Heterogeneous Mixture –

14. Distillation -

15. Filtration –

16. Chromatography -

17. Crystallization -

18. Reactants -

19. Products -

20. Chemical Equation -

21. Law of Conservation of Mass -

23. A 3.50-g sample of pure copper reacts with chlorine to form 8.76 g of a copper chloride compound.

How many grams of chlorine are used in this reaction? (show work and give units) **5.26 g**

22. Identify the following as quantitative or qualitative data...

Color - **qualitative**

Volume - **quantitative**

Pressure - **quantitative**

Shape - **qualitative**

Temperature - **quantitative**

Odor - **qualitative**

24. Identify the following as a chemical or physical property...

Remains unchanged in oxygen - **chemical**

Photosynthesis - **chemical**

Liquid at room temperature - **physical**

Burning coal – **chemical**

Blue - **physical**

Snapping a branch off a tree – **physical**

Conducts electricity - **chemical**

Water boiling – **physical**

Changes colors when heated - **chemical**

Chrome tarnishing – **chemical**

Grinding cocoa beans – **physical**

Baking a cake - **chemical**

25. Identify the following as heterogeneous or homogeneous mixtures

Blood - **heterogeneous**

Gasoline - **homogeneous**

Limestone - **heterogeneous**

Baking soda - **homogeneous**

Polluted air - **heterogeneous**

Top soil - **heterogeneous**

Glass - **homogeneous**

Grape juice - **homogeneous**

Lotion - **homogeneous**

Milk - **homogeneous**

26. Identify the following as an element or a compound

Copper - **element**

Carbon Monoxide - **compound**

Vinegar - **compound**

Bronze - **compound**

Oxygen - **element**

Plutonium - **element**