**Write half-reactions for each of the following redox reactions. Identify each half-reaction as being either oxidation or reduction.**

1. SnS2 + O2 → SnO2 + SO2
2. Mg + N2 → Mg3N2
3. Al + Cl2 → AlCl3
4. NH3 + PbO → N2 + Pb + H2O
5. I2 + Na2S2O3 → Na2S2O4 + NaI
6. Sn + 2HCl → SnCl2 + H2

**Write half-reactions for each of the following redox reactions. Identify each half-reaction as either oxidation or reduction.**

1. SnS2 + O2 → SnO2 + SO2

S2- → S4+ + 6e- oxidation

O2 + 4e- → 2O2- reduction

1. Mg + N2 → Mg3N2

Mg → Mg2+ + 2e- oxidation

N2 + 6e- → 2N3- reduction

1. Al + Cl2 → AlCl3

Cl2 + 2e- → 2Cl- reduction

1. NH3 + PbO → N2 + Pb + H2O
2. I2 + Na2S2O3 → Na2S2O4 + NaI

S2+ → S3+ + e- oxidation

1. Sn + 2HCl → SnCl2 + H2